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This reaction follows the equation:  $N_2 + 3H_2 \rightarrow 2NH_3$

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$\Delta G^0 = \Delta H^0 - T\Delta S^0 = 192 \text{ kJ mol}^{-1} - (298 \text{ K})(0.192 \text{ kJ mol}^{-1} \text{ K}^{-1}) = 35 \text{ kJ mol}^{-1}$  Alternatively, one can also find the temperature for which  $\Delta G = 0$ ,  $T = \Delta H^0 / \Delta S^0$ .

Introduction to Chemical Engineering: Chemical Reaction ...

One of the most difficult problems in chemical reaction engineering is to navigate in a reaction network efficiently in order to optimize the production of the desired target component (Problem 4). In this field again catalysis is a main tool.

1 Basic Problems of Chemical Reaction Engineering and ...

example problems before going on to other solutions. P2-5. This is a reasonably challenging problem that reinforces Levenspiel's plots. P2-6. Straight forward problem alternative to problems 7, 8, and 11. P2-7. To be used in those courses emphasizing bio reaction engineering. P2-8. The answer gives ridiculously large reactor volume.

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